

Chapter 02

The Chemical Basis of Life I: Atoms, Molecules, and Water

Multiple Choice Questions

1. The atomic number of an atom is
- A. the number of protons in the atom.
 - B. the number of neutrons in the atom.
 - C. the number of protons and electrons in the atom.
 - D. the number of protons and neutrons in the atom.
 - E. None of these choices are correct.

Blooms Level: 1. Remember

LO: 02.01.03 Relate atomic structure to the periodic table of the elements.

Section: 02.01 Atoms

Topic: Chemistry

2. The smallest functional unit of living organisms is
- A. atoms
 - B. molecules
 - C. proteins
 - D. water
 - E. salt

Blooms Level: 1. Remember

General LO: Compare relative scales of biological structures and processes.

Section: 02.01 Atoms

Topic: General

Check All That Apply Questions

3. With an atomic mass of 16 and an atomic number of 8, it follows that oxygen
__X__ has eight electrons.
_____ has 16 neutrons.
__X__ can readily form bonds with 2 other atoms.
_____ weighs 16 grams.

Blooms Level: 2. Understand

LO: 02.01.03 Relate atomic structure to the periodic table of the elements.

Section: 02.01 Atoms

Topic: Chemistry

Multiple Choice Questions

4. The nucleus of an atom is composed of
A. protons.
B. neutrons.
C. electrons.
D. protons and neutrons.
E. protons and electrons.

Blooms Level: 1. Remember

LO: 02.01.01 Understand the general structure of atoms and their constituent particles.

Section: 02.01 Atoms

Topic: Chemistry

5. Ernest Rutherford's key experiment on alpha particle bombardment of gold foil was important to the development of
- A. detection methods for protons.
 - B. alpha particle emitters.
 - C. gold as an element.
 - D.** the modern model for atomic structure.
 - E. the concept that atoms have a homogenous distribution of protons throughout the atom.

Blooms Level: 1. Remember

LO: 02.01.01 Understand the general structure of atoms and their constituent particles.

Section: 02.01 Atoms

Topic: Chemistry

6. If a scientist were to shoot protons through an atom as Rutherford did with gold foil, he or she would likely find that
- A. most of the protons passed straight through the atom.
 - B. few of the protons passed straight through the atom.
 - C. most of the protons deflected or bounced back from the atom.
 - D.** most of the protons would be absorbed by the atom.
 - E. the atom would emit protons.

Blooms Level: 3. Apply

LO: 02.01.01 Understand the general structure of atoms and their constituent particles.

Section: 02.01 Atoms

Topic: Chemistry

7. The first, inner-most energy shell of an atom
- A. can have a maximum of 8 electrons.
 - B.** can have a maximum of 2 electrons.
 - C. is called the 2p orbital.
 - D. is called the 1s orbital and can have a maximum of 8 electrons.
 - E. is called the 2p orbital and can have a maximum of 2 electrons.

Blooms Level: 1. Remember

LO: 02.01.02 Discuss the way electrons orbit the nucleus of an atom within discrete energy levels.

Section: 02.01 Atoms

Topic: Chemistry

8. Tritiated hydrogen (^3H) differs from hydrogen (^1H) in that
- A. ^3H has 2 more protons than ^1H .
 - B. ^3H has 2 more electrons than ^1H .
 - C.** ^3H has 2 more neutrons than ^1H .
 - D. ^3H has the same number of neutrons as ^1H .
 - E. ^3H has a different electron configuration than ^1H .

Blooms Level: 2. Understand

LO: 02.01.05 Explain how a single element may exist in more than one form, called isotopes, and how certain isotopes have importance in human medicine.

Section: 02.01 Atoms

Topic: Chemistry

9. Isotopes are different forms of the same element that
- A. differ in their number of neutrons.
 - B. differ in their number of protons.
 - C. are all produced artificially.
 - D. cannot form covalent bonds.
 - E. cannot form ions.

Blooms Level: 2. Understand

LO: 02.01.05 Explain how a single element may exist in more than one form, called isotopes, and how certain isotopes have importance in human medicine.

Section: 02.01 Atoms

Topic: Chemistry

10. The element found in most abundance in living organisms is
- A. calcium.
 - B. iron.
 - C. iodine.
 - D. hydrogen.
 - E. sodium.

Blooms Level: 2. Understand

LO: 02.01.06 Know which elements make up most of the mass of all living organisms.

Section: 02.01 Atoms

Topic: Chemistry

11. Nitrogen has 7 electrons and can form a maximum of _____ bonds with other elements.

- A. 1
- B. 2
- C. 3**
- D. 4
- E. 5

Blooms Level: 2. Understand

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

12. Molecules

- A. are derived from the ionic bonding of two or more atoms.
- B. have the same physical properties as the atoms from which they were derived.
- C. are not important in biological processes.
- D. can form from the covalent bonding of two or more atoms.**
- E. cannot have a charge.

Blooms Level: 2. Understand

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

13. Identify the ion from below.

- A. Ca^{2+}
- B. He
- C. H_2
- D. CO_2
- E. KCl^-

Blooms Level: 2. Understand

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

14. Carbon has 4 electrons and hydrogen has 1 electron in its outermost electron shell. A carbon atom can form covalent bonds with how many hydrogen atoms?

- A. 0
- B. 1
- C. 2
- D. 3
- E. 4

Blooms Level: 3. Apply

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

15. When one atom loses an electron to another atom, it results in the formation of
- A. a polar covalent bond and a new molecule.
 - B.** cations and anions that can form ionic bonds.
 - C. a covalent bond between the two.
 - D. many hydrogen bonds.
 - E. a nonpolar covalent bond that is difficult to break.

Blooms Level: 2. Understand

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

16. The strongest chemical bonds are
- A. hydrogen bonds.
 - B. Van der Waal forces.
 - C. hydrophobic interactions.
 - D. ionic bonds.
 - E.** covalent bonds.

Blooms Level: 1. Remember

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

17. What type of bonding is likely to occur between two water molecules or strands of DNA?

- A. covalent
- B. ionic
- C. hydrogen**
- D. both hydrogen and covalent
- E. both hydrogen and ionic

Blooms Level: 2. Understand

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

18. Carbon and hydrogen have similar electronegativities and combine together to form hydrocarbon molecules. What type of bonds form between these atoms?

- A. hydrogen
- B. ionic
- C. polar covalent
- D. nonpolar covalent**
- E. electrostatic

Blooms Level: 2. Understand

LO: 02.02.02 Explain the concept of electronegativity and how it contributes to the formation of polar and nonpolar covalent bonds.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

19. What type of bonds form from the unequal sharing of electrons?

- A. hydrogen
- B. ionic
- C. polar covalent**
- D. nonpolar covalent
- E. electrostatic

Blooms Level: 2. Understand

LO: 02.02.02 Explain the concept of electronegativity and how it contributes to the formation of polar and nonpolar covalent bonds.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

20. In water, MgCl_2 dissociates into Mg^{2+} and Cl^- . Based on this information what type of bond is involved in the formation of MgCl_2 ?

- A. hydrogen
- B. ionic**
- C. polar covalent
- D. nonpolar covalent
- E. electrostatic

Blooms Level: 3. Apply

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

21. When one oxygen atom shares two pairs of electrons with another oxygen atom, O₂ is formed via a(n)

- A. single covalent bond.
- B. double covalent bond.**
- C. triple covalent bond.
- D. ionic bond.
- E. hydrogen bond.

Blooms Level: 2. Understand

LO: 02.02.02 Explain the concept of electronegativity and how it contributes to the formation of polar and nonpolar covalent bonds.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

22. The LEAST hydrophilic substance is

- A. salt.
- B. an ion.
- C. oil.**
- D. an amphipathic molecule.
- E. a gas.

Blooms Level: 2. Understand

LO: 02.03.02 List the properties of water that make it a valuable solvent, and distinguish between hydrophilic and hydrophobic substances.

Section: 02.03 Properties of Water

Topic: Chemistry

23. Amphipathic molecules

- A. possess only hydrophilic properties.
- B. possess only hydrophobic properties.
- C.** possess both hydrophilic and hydrophobic properties.
- D. possess neither hydrophilic nor hydrophobic properties.
- E. tend not to interact with other molecules.

Blooms Level: 1. Remember

LO: 02.03.02 List the properties of water that make it a valuable solvent, and distinguish between hydrophilic and hydrophobic substances.

Section: 02.03 Properties of Water

Topic: Chemistry

24. For water to vaporize

- A. energy must be supplied.
- B. energy must be released.
- C. hydrogen bonds are broken.
- D.** both energy must be supplied and hydrogen bonds broken.
- E. both energy must be released and hydrogen bonds broken.

Blooms Level: 2. Understand

LO: 02.03.01 Describe how hydrogen bonding determines many properties of water.

Section: 02.03 Properties of Water

Topic: Chemistry

25. The molarity of a solution is
- A. a measure of solute concentration.
 - B. the weight of a solid substance.
 - C. often expressed as grams per unit volume.
 - D. reflects a measure of the amount of oil dissolved in water.
 - E. a scientific term for determining the solubility of a substance in water.

Blooms Level: 1. Remember

LO: 02.03.03 Understand how the molarity of a solution the number of moles of a solute per liter of solution is used to measure the concentration of solutes in solution.

Section: 02.03 Properties of Water

Topic: Chemistry

26. Based on the colligative properties of water, what would happen if one were to add a solute to water?
- A. The freezing point of water would decrease.
 - B. The freezing point of water would increase.
 - C. The boiling point of water would increase.
 - D. Both the freezing point of water would decrease and the boiling point of water would increase.
 - E. Nothing would change with respect to the freezing point or boiling point of water.

Blooms Level: 3. Apply

LO: 02.03.02 List the properties of water that make it a valuable solvent, and distinguish between hydrophilic and hydrophobic substances.

Section: 02.03 Properties of Water

Topic: Chemistry

27. Water

- A. is nonpolar.
- B. has a low heat of vaporization.
- C.** has cohesive properties.
- D. evaporates and increases body temperature.
- E. is a relatively poor solvent.

Blooms Level: 2. Understand

LO: 02.03.02 List the properties of water that make it a valuable solvent, and distinguish between hydrophilic and hydrophobic substances.

Section: 02.03 Properties of Water

Topic: Chemistry

28. If orange juice has a pH of 4 then it can be described as

- A. having a H^+ concentration is 4.
- B.** an acidic solution.
- C. an alkaline solution.
- D. an acidic solution with a H^+ concentration of 4.
- E. None of these choices are correct.

Blooms Level: 3. Apply

LO: 02.03.05 Explain how water has the ability to ionize into hydroxide ions and into hydrogen ions and how the H concentration is expressed as a solution's pH.

Section: 02.03 Properties of Water

Topic: Chemistry

29. The most significant role played by pH buffers is to
- A. prevent fluctuations in the acidity of solutions.
 - B. increase the strength of acids and bases.
 - C. prevent fluctuations in the salinity of solutions.
 - D.** limit major shifts in the amount of H^+ and OH^- in solution.
 - E. keep pH low.

Blooms Level: 5. Evaluate

LO: 02.03.06 Give examples of how buffers maintain a stable environment in an animal's body fluids.

Section: 02.03 Properties of Water

Topic: Chemistry

30. One of the ways a pH buffer helps to maintains homeostasis is by
- A. increasing the amount of H^+ in an acidic solution.

B.

reducing the amount of H^+ in an acidic solution.

- C. reducing the amount of H^+ in an alkaline solution.
- D. increasing the amount of OH^- ions in an alkaline solution.
- E. reducing the amount of OH^- in an acidic solution.

Blooms Level: 3. Apply

LO: 02.03.05 Explain how water has the ability to ionize into hydroxide ions and into hydrogen ions and how the H concentration is expressed as a solution's pH.

Section: 02.03 Properties of Water

Topic: Chemistry

31. The addition of a strong acid like HCl to an aqueous solution would result in
- A. the release of H^+ into the solution.
 - B. an increase in pH.
 - C. a decrease in pH.
 - D. both the release of H^+ and an increase in pH.
 - E.** both the release of H^+ and a decrease in pH.

Blooms Level: 4. Analyze

LO: 02.03.05 Explain how water has the ability to ionize into hydroxide ions and into hydrogen ions and how the H concentration is expressed as a solution's pH.

Section: 02.03 Properties of Water

Topic: Chemistry

True / False Questions

32. One gram of hydrogen, which has an atomic mass of 1, would have fewer atoms than 1 gram of carbon that has an atomic mass of 12.

FALSE

Blooms Level: 3. Apply

LO: 02.01.03 Relate atomic structure to the periodic table of the elements.

Section: 02.01 Atoms

Topic: Chemistry

33. Isotopes are different forms of the same element.

TRUE

Blooms Level: 2. Understand

LO: 02.01.05 Explain how a single element may exist in more than one form, called isotopes, and how certain isotopes have importance in human medicine.

Section: 02.01 Atoms

Topic: Chemistry

34. Sulfur 35 (^{35}S) is an isotope of ^{32}S . These elements differ in their number of neutrons.

TRUE

Blooms Level: 2. Understand

LO: 02.01.05 Explain how a single element may exist in more than one form, called isotopes, and how certain isotopes have importance in human medicine.

Section: 02.01 Atoms

Topic: Chemistry

35. Helium is an inert gas that rarely reacts with other elements because it has the maximum number of valence electrons in its outer shell.

TRUE

Blooms Level: 2. Understand

LO: 02.01.01 Understand the general structure of atoms and their constituent particles.

Section: 02.01 Atoms

Topic: Chemistry

36. If lithium has an atomic number of 3 then it will have 1 valence electron.

TRUE

Blooms Level: 2. Understand

LO: 02.01.03 Relate atomic structure to the periodic table of the elements.

Section: 02.01 Atoms

Topic: Chemistry

37. The electronegativity of an atom is a measure of its ability to attract electrons to its outer shell from another atom.

TRUE

Blooms Level: 2. Understand

LO: 02.02.02 Explain the concept of electronegativity and how it contributes to the formation of polar and nonpolar covalent bonds.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

38. Table salt forms from sodium and chloride via hydrogen bonding.

FALSE

Blooms Level: 2. Understand

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

39. Molecules are generally rigid structures and rarely change shape.

FALSE

Blooms Level: 2. Understand

LO: 02.02.03 Describe how a molecule's shape is important for its ability to interact with other molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

40. The presence of salt helps prevent oceans from freezing.

TRUE

Blooms Level: 2. Understand

LO: 02.03.02 List the properties of water that make it a valuable solvent, and distinguish between hydrophilic and hydrophobic substances.

Section: 02.03 Properties of Water

Topic: Chemistry

41. A dehydration reaction that builds larger molecules from smaller units requires the addition of a water molecule.

FALSE

Blooms Level: 2. Understand

LO: 02.02.04 Understand the concepts of a chemical reaction and chemical equilibrium.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

42. The hydroxyl (OH^-) concentration of a solution with a pH of 8 would be 10^{-6} molar.

TRUE

Blooms Level: 3. Apply

LO: 02.03.05 Explain how water has the ability to ionize into hydroxide ions and into hydrogen ions and how the H concentration is expressed as a solution's pH.

Section: 02.03 Properties of Water

Topic: Chemistry

43. Most enzymes or bioactive molecules work effectively within a broad range of pH.

FALSE

Blooms Level: 2. Understand

General LO: Compare relative scales of biological structures and processes.

Section: 02.02 Chemical Bonds and Molecules

Topic: Molecular Biology

Multiple Choice Questions

44. $\text{Zn} + 2\text{H}^+ = \text{Zn}^{2+} + \text{H}_2$ is an example of a redox reaction. What is happening during this interaction? Is a bond created between the atoms during this reaction?

- A. Oxidation reaction and acceptance of an electron; bond is formed.
- B. Reduction reaction and acceptance of an electron; no bond is formed.**
- C. Reduction reaction and donation of an electron; no bond is formed.
- D. Covalent interaction; bond is formed.

Blooms Level: 4. Analyze

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Type: Quantitative Reasoning

45. You notice that the majority of the electrons in NaCl spend their time around the chlorine. You also notice that the electrons in H₂ are evenly distributed among the two atoms. Which two types of bonds are represented in these molecules?

- A. Covalent bonds in NaCl; ionic bonds in H₂.
- B. Covalent bonds in NaCl; covalent bonds in H₂.
- C. Ionic bonds in NaCl; ionic bonds in H₂.
- D.** Ionic bonds in NaCl; covalent bonds in H₂.

Blooms Level: 4. Analyze

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

46. A bottle of Na in solution and a bottle of Cl in solution are mixed together. What type of bond will be created between the atoms, and what will be the product?

- A. Covalent bonds; sodium chlorine
- B.** Ionic bonds; table salt
- C. Hydrogen bonds; sodium hydroxide
- D. Carbon bonds; carboxyl groups

Blooms Level: 3. Apply

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

47. You've been asked to stabilize a compound whose general state is altered by excess electrons. The element you would add to the compound to most effectively stabilize the compound would be? Why?

- A. Carbon, because it is capable of neutralizing electrons.
- B. Nitrogen, because it have five electrons on its outer shell.
- C.** Fluorine, because it is the greediest atom on the periodic table.
- D. Oxygen, because it can easily bind with the compound.

Blooms Level: 3. Apply

LO: 02.01.03 Relate atomic structure to the periodic table of the elements.

Section: 02.01 Atoms

Topic: Chemistry

48. You want to simulate the production of carbon dioxide (CO_2) in a laboratory setting using carbon and oxygen atoms. What type of reactions do you need to facilitate in order to create CO_2 ?

- A. An oxidation, or the gain of an electron, and a reduction, or the loss of an electron.
- B. An oxidation, or the loss and electron, and a reduction, or the loss of an electron.
- C. An oxidation, or the gain of an electron, and a reduction, or the gain of electron.
- D.** An oxidation, or the loss of an electron, and a reduction, or the gain of an electron.

Blooms Level: 3. Apply

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

49. Five unknown compounds are added to water. Four of the compounds go into solution while one does not. What property does water possess that allows these four compounds to dissolve? Why might the fifth compound not dissolve?

- A.** The positive and negative charge in water will dissolve many substances; the substance is not structurally similar to water.
- B. The negative charge of water dissolves many substances; the substance is too structurally similar to water.
- C. The positive charge of water dissolves many substances; the substance is too structurally similar to water.
- D. The nonpolar qualities of water dissolves many substances; the substance is not structurally similar to water.

Blooms Level: 3. Apply

LO: 02.03.01 Describe how hydrogen bonding determines many properties of water.

Section: 02.03 Properties of Water

Topic: Chemistry

50. 1 mole = 1000 millimoles (mmol); 1 millimole = 1000 micromoles (μmol). If a solution contains 38231 μmol , what is that amount in mmol?

- A. 382.31 mmol
- B.** 38.231 mmol
- C. 3.8231 mmol
- D. 3823.1 mmol

Blooms Level: 3. Apply

LO: 02.03.03 Understand how the molarity of a solution the number of moles of a solute per liter of solution is used to measure the concentration of solutes in solution.

Section: 02.03 Properties of Water

Type: Quantitative Reasoning

51. If 1000 millimoles make up a mole, how many grams of magnesium (Mg), which has an atomic mass of 24.305, will make a solution of 150 μmol ?

- A. 3.6mg
- B. 2.4mg
- C. 0.24mg
- D. 36mg

Blooms Level: 3. Apply

LO: 02.01.04 Quantify atomic mass using units such as daltons and moles.

Section: 02.01 Atoms

Type: Quantitative Reasoning

52. Using the periodic table as your tool, identify the atomic characteristic that would most quickly and efficiently identify any single element:

- A. number of shells
- B. number of neutrons
- C. number of protons and electrons
- D. number of neutrons and electrons

Blooms Level: 5. Evaluate

LO: 02.01.03 Relate atomic structure to the periodic table of the elements.

Section: 02.01 Atoms

Type: Quantitative Reasoning

53. In the periodic table, the value that refers to the number of protons and neutrons is:

- A. atomic mass.
- B. molecular molarity.
- C. atomic molarity.
- D. molecular number.

Blooms Level: 5. Evaluate

LO: 02.01.03 Relate atomic structure to the periodic table of the elements.

Section: 02.01 Atoms

Topic: Chemistry

54. You've been given three new elements. One element had all its protons removed, one element had all its neutrons removed, and one element had all its electrons removed. The effect that would have the largest effect on atomic mass would be

- A. Removing the protons.
- B. Removing the neutrons.**
- C. Removing the electrons.
- D. All the changes would affect the atomic mass.

Blooms Level: 3. Apply

LO: 02.01.01 Understand the general structure of atoms and their constituent particles.

Section: 02.01 Atoms

Topic: Chemistry

55. You have been asked to synthesize a new isotope for cadmium. Which part of the original atom would you need to manipulate in order to create an isotope?

- A. Neutrons**
- B. Protons
- C. Protons and neutrons
- D. Electrons

Blooms Level: 3. Apply

LO: 02.01.05 Explain how a single element may exist in more than one form, called isotopes, and how certain isotopes have importance in human medicine.

Section: 02.01 Atoms

Topic: Chemistry

56. The single atom you would choose to remove from living organisms in order to remove the highest percentage of atoms would be:

- A. Oxygen
- B. Nitrogen
- C. Hydrogen
- D. Carbon

Blooms Level: 3. Apply

LO: 02.03.04 Discuss the properties of water that are critical for the survival of living organisms.

Section: 02.03 Properties of Water

Topic: Chemistry

57. You have been given vials of H_2 , Na, H_2O , Hg, and CH_4 . What are the majority of your vials filled with?

- A. Liquids at room temperature.
- B. Gases
- C. Molecules
- D. Carboxyls

Blooms Level: 3. Apply

LO: 02.02.01 Compare and contrast the types of atomic interactions that lead to the formation of molecules.

Section: 02.02 Chemical Bonds and Molecules

Topic: Chemistry

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58. Water has fewer hydrogen atoms than lemon juice and a pH of around 7. Predict what will happen to the pH level of water when lemon juice is added.

- A. The pH will become higher.
- B.** The pH will become lower.
- C. The pH will remain the same.
- D. There is not enough information to decide.

Blooms Level: 3. Apply

LO: 02.03.05 Explain how water has the ability to ionize into hydroxide ions and into hydrogen ions and how the H concentration is expressed as a solution's pH.

Section: 02.03 Properties of Water

Topic: Chemistry

59. This statement is true when comparing solutions with a pH of 6 and a pH of 8.

- A. The solution with a pH of 8 has a 100 times higher concentration of hydrogen ions than a solution with a pH of 6.
- B. Solutions with a pH of 8 has a 2 times lower concentration of hydrogen ions than a solution with a pH of 6.
- C.** The solution with a pH of 6 has a 100 times higher concentration of hydrogen ions than a solution with a pH of 8.
- D. The solution with a pH of 6 has a 100 times lower concentration of hydrogen ions than a solution with a pH of 8.
- E. The solution with a pH of 6 has a 2 times higher concentration of hydrogen ions than a solution with a pH of 8.

Blooms Level: 4. Analyze

LO: 02.03.05 Explain how water has the ability to ionize into hydroxide ions and into hydrogen ions and how the H concentration is expressed as a solution's pH.

Section: 02.03 Properties of Water

Topic: Chemistry